

## Online Library Chem Activity 42 Answers

# Chem Activity 42 Answers

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6. a)  $\text{NH}_4^+$  + b)  $\text{OH}^-$  - c) A conjugate acid-base pair are two chemical species that differ by a proton. ChemActivity 42

Answers to Problems and Exercises 1.

$\text{SO}_4^{2-}$  ;  $\text{CO}_3^{2-}$  ;  $\text{OH}^-$  ;  $\text{O}^{2-}$  ;  $\text{H}_2\text{O}$  ;

$\text{NH}_3$  ;  $\text{CH}_3\text{NH}_2$  ;  $\text{F}^-$  ;  $\text{CH}_3\text{COO}^-$ . 2.

$\text{HSO}_4^-$  ;  $\text{HCO}_3^-$  ;  $\text{H}_3\text{O}^+$  ;  $\text{H}_2\text{O}$  ;  $\text{OH}^-$  ;

$\text{NH}_4^+$  ;  $\text{CH}_3\text{NH}_3^+$  ;  $\text{HCN}$  ;  $\text{CH}_3\text{COOH}$

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; HF ; H<sub>2</sub>CO<sub>3</sub> ; NH<sub>3</sub> .

## **ANSWERS\_to\_CTQs\_and\_Exercises\_Chemactivities\_42\_43 ...**

4. The bond energy (in MJ/mole) is the energy required to break one mole of the specified bonds. Triple bonds share six electrons between two atoms and are stronger (have greater bond energy)

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than double bonds, which share four electrons between two atoms.

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using the POGIL methodology. This text follows the principles of inquiry-based learning and correspondingly emphasizes underlying chemistry concepts and the reasoning behind them.

**Chemistry: A Guided Inquiry, 7th Edition | Wiley**

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PLEASE NOTE: If you have a question about these answers, it is your responsibility to come to office hours or ask during class work time. UNIT 12 - HW Practice Keys - ChemActivity 51: Cell Voltage - ChemActivity 50: Electrochemical Cell UNIT 11 - HW Practice Keys - ChemQuest 55: Free Energy - ChemQuest 54: 2nd Law of

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Thermodynamics

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Answer Keys, Study Guides and Vocabulary Sets. CHEMISTRY is the study of matter, its properties, how and why substances combine or separate to form other substances, and how substances interact with energy. The five main branches of chemistry include analytical chemistry, physical chemistry, organic chemistry, inorganic chemistry

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and biochemistry.

## **Printable Chemistry Worksheets and Answer Keys, Study ...**

ChemActivity 42 : Sep 21: Exam # 1 :

Sep 22: Details of acid and base  
chemistry: ChemActivity 43

ChemActivity 44 ChemActivity 45: CA 42

- 45: Sep 23: Solving weak acid base

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problems: ChemActivity 45: Sep 26:  
Details of strong acid/base reactions:  
ChemActivity 45(a) CA 45(a) Sep 28: pH  
of salt solutions : ChemActivity 47:  
ChemActivity 47: Sep 30

## **Chem 124 Homework - Wofford College**

To the Instructor We are very pleased to

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## Answers

be able to provide these selected activities from Chemistry: A Guided Inquiry, 5th Edition for use in high school classrooms - particularly for the new AP Chemistry Curriculum. Our main goal is to provide materials that might be helpful in

**Chemistry: A Guided Inquiry 5e**

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## Answers

Types of Chemical Reactions Do atoms rearrange in predictable patterns during chemical reactions? why? Recognizing patterns allows us to predict future behavior. rWearher experrs use pafterns to predicr danger- ous storms so people can get iheir families to safery Political analysrs use patterns to prediit elecrrion out- comes. Similarly, chemists classify

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## Answers

chemical equations according to their ...

### **Chemical - LTHS Answers**

4. The answer here depends on the students' answer to CTQ 3. 5.a) The IE 1 for He is greater than the IE 1 for H because the nuclear charge on a He atom is +2 whereas the nuclear charge on H is +1. b) The IE 1 for Li is less than

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the IE 1 for He because at least one of the electrons of Li must be farther away from the nucleus than any electron ...

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He has a link to Chemistry-11 and to Chemistry -12. The ChemTeam Check out the gallery of chemistry photos, the



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classic papers, the tutorials, and comprehensive collection of old NChO exams with answers for practice.

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ChemActivity 42 Know the Bronsted-Lowery definition of an acid and a base. In a chemical reaction identify the

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Bronsted acid and base in the reactants.  
Determine the conjugate base of a given acid and determine the conjugate acid of a base.

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For example:  $94\text{Be} + 42\text{He} + \text{Ion } 14\text{N}$   
 $+ 42\text{He } 1780 + 1\text{H } 21\ 27\ 4521\text{Sc} + \text{Ion}$   
 $4219\text{K} + 42\text{He}$  In nuclear equations, the total number of positive charges

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(represented by the atomic numbers) of the reactants (substances on the left of the reaction arrow) equals the total number of positive charges of the products (substances on the right of the reaction arrow).

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