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## **Cation Group Ii Copper Arsenic**

Cation Group II (Copper Arsenic group )

Group IIA:  $\text{Hg}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Bi}^{3+}$ ,  $\text{Cu}^{2+}$ ,  
 $\text{Cd}^{2+}$  Group IIB:  $\text{Sn}^{2+}$ ,  $\text{Sb}^{3+}$ ,  $\text{As}^{3+}$  This analytical group of cations is composed of eight cations which are subdivided into the copper group consisting of mercuric, bismuth, cadmium, copper and lead; and the arsenic group

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consisting of arsenic, antimony and tin.

### **Cation Group II (Copper Arsenic group )**

Cations of group II are known as Copper arsenic group in qualitative analysis.

This group II has following cations.

Group II A: Hg +2, Pb +2, Bi +3, Cu +2, Cd +2. Group II B: Sn +2, Sb +3, As +3.

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Cations of whole group II is precipitated as Metal sulphide by adding  $H_2S$  in acidic medium.

**Answered: Explain the theoretical aspect of the... | bartleby**

Methods for the identification of mercury, bismuth, copper, cadmium, arsenic, antimony, and tin. Specific tests



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for cations of group II | Journal of  
Chemical Education ACS

## **Specific tests for cations of group II | Journal of ...**

Group 2 cations - Insoluble sulfides in  
acidic medium  $\square$  HgS, PbS, CuS, Bi<sub>2</sub>S<sub>3</sub>,  
CdS, As<sub>2</sub>S<sub>3</sub>, Sb<sub>2</sub>S<sub>3</sub>, SnS<sub>2</sub> Copper  
subgroup - Hg<sup>2+</sup>, Pb<sup>2+</sup>, Cu<sup>2+</sup>, Bi<sup>3+</sup>,

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$\text{Cd}^{2+}$  The sulfides are insoluble in KOH solution, only soluble in nitric acid

Arsenic subgroup -  $\text{As}^{3+}$ ,  $\text{Sb}^{3+}$ ,  $\text{Sn}^{4+}$

The sulfides are thioamphoteric that are soluble in  $\text{KOH}(\text{aq})$  and nitric acid

### **Qualitative Analysis of Group II Cations**

Group II cation Page 64 Cation Group II

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(Copper Arsenic group ) Group IIA: Hg 2+, Pb 2+, Bi 3+, Cu 2+, Cd 2+ Group IIB: Sn 2+, Sb 3+, As 3+ This analytical group of cations is composed of eight cations which are subdivided into the copper group consisting of mercuric, bismuth, cadmium, copper and lead; and the arsenic group consisting of arsenic, antimony and tin. The group reagent

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(precipitating agent) of cation group II is hydrogen sulphide in acid medium (0.3 M HCl).

**GRP 2 cations - Cation Group II(Copper Arsenic group Group ...**  
Cation analysis for Arsenic ions in lab class 11 and 12 - Duration: 4:34. Seema Makhijani 7,881 views. 4:34. Cation Test:

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Copper(II) Ions - Duration: 3:32 ... FLAME  
TEST group 5 salt ...

### **Cation analysis for Copper ions in lab class 11 and 12**

Group II. Cations do not react with hydrochloric acid, but form precipitates with hydrogen sulphide in dilute mineral acid medium. Ions of this group are: - II

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A. Mercury (II), Copper, Bismuth, Cadmium Sulphides of II A are insoluble in  $(\text{NH}_4)_2\text{S}$ . IIB. Arsenic (III), Arsenic (V), Antimony (III), Antimony (V), Tin (II) and Tin (IV).

### **Experiment # 9 : Cation Analysis: Group ( I )**

The elements of the copper-arsenic

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group are  $\text{Hg}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Bi}^{3+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Cd}^{2+}$ ,  $\text{As}^{3+}$ ,  $\text{Sb}^{3+}$  and  $\text{Sn}^{4+}$ . They form sulphides that are insoluble in dilute  $\text{HCl}$ . It is not the case for sulphides of the following groups so we can isolate the ions from the copper-arsenic group from the ions of the other groups.

### **Chapter 10b: The Copper-Arsenic**

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## **group - BORZUYA UNIVERSITY**

Group II Cations ( $\text{Hg}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Bi}^{3+}$ ,  $\text{Cd}^{2+}$ ,  $\text{As}^{3+}$ ,  $\text{Sb}^{3+}$  and  $\text{Sn}^{4+}$  - insoluble sulphides in acidic medium):  
After the insoluble chlorides are isolated, the pH of the solution is adjusted to 0.5 and then  $\text{H}_2\text{S}$  is added.

## **Systematic Analysis of Cations -**



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## **Chemistry Practicals Class 12**

IIA group consists of mercury(II)  $\text{Hg}^{2+}$ , lead(II)  $\text{Pb}^{2+}$ , bismuth(III)  $\text{Bi}^{3+}$ , copper(II)  $\text{Cu}^{2+}$ , cadmium(II)  $\text{Cd}^{2+}$  and they are insoluble in ammonium polysulphide. IIB group consists of arsenic(III)  $\text{As}^{3+}$ , arsenic(V)  $\text{As}^{5+}$ , antimony(III)  $\text{Sb}^{3+}$ , antimony(V)  $\text{Sb}^{5+}$ , tin(II)  $\text{Sn}^{2+}$  and tin(IV)  $\text{Sn}^{4+}$  and these

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are soluble in ammonium polysulphide.

### **QUALITATIVE ANALYSIS OF METAL CATIONS**

Definitions. Action level means a concentration of inorganic arsenic of 5 micrograms per cubic meter of air ( $5 \mu\text{g}/\text{m}^3$ ) averaged over any eight (8) hour period. Assistant Secretary means

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the Assistant Secretary of Labor for Occupational Safety and Health, U.S. Department of Labor, or designee. Authorized person means any person specifically authorized by the employer whose duties require the ...

**1910.1018 - Inorganic arsenic. | Occupational Safety and ...**

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It is mainly found in minerals. The most common arsenic-containing mineral is arsenopyrite. Others include realgar, orpiment and enargite. Most arsenic is produced as a by-product of copper and lead refining. It can be obtained from arsenopyrite by heating, causing the arsenic to sublime and leave behind iron(II) sulfide.

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## **Arsenic - Element information, properties and uses ...**

The present study aims to evaluate the competitive biosorption of lead, cadmium, copper, and arsenic ions by using native algae. A series of experiments were carried out in a batch reactor to obtain equilibrium data for

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adsorption of single, binary, ternary, and quaternary metal solutions.

### **Competitive biosorption of lead, cadmium, copper, and ...**

Arsenic is a chemical element with the symbol As and atomic number 33.

Arsenic occurs in many minerals, usually in combination with sulfur and metals,

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but also as a pure elemental crystal. Arsenic is a metalloid. It has various allotropes, but only the gray form, which has a metallic appearance, is important to industry.. The primary use of arsenic is in alloys of lead (for example, in car ...

**Arsenic - Wikipedia**

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Group I Cation Page 42 Group Group  
reagent ions Formula of precipitate  
Group 1 (silver group) cold, dilute-  
hydrochloric acid  $Hg^{2+}$ ,  $Pb^{2+}$ ,  $Ag^{+}$   
 $Hg_2Cl_2$ ,  $PbCl_2$ ,  $AgCl$  Group II (Copper  
and arsenic group)  $H_2S$  in presence of  
dilute HCl  $Hg^{2+}$ ,  $Pb^{2+}$ ,  $Bi^{3+}$ ,  $Cu^{2+}$ ,  
 $Cd^{2+}$ ,  $Sn^{2+}$ ,  $Sb^{3+}$ ,  $As^{3+}$   $HgS$ ,  $PbS$ ,  $Bi_2S_3$ ,  
 $CuS$ ,  $CdS$ ,  $SnS$ ,  $Sb_2S_3$ ,  $As_2S_3$



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Group III

## **Analytical Separation of Cations**

Table of Contents  
Roasting of Copper-  
Ores  
Smelting Calcined Ores  
Calcining Coarse Metal  
Smelting Calcined Coarse  
Metal  
Roasting White Metal to Blister-  
Copper  
Refining Blister-  
Copper  
Bessemerizing Copper

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RegulusPyritic SmeltingElimination by  
Wet Processes of ExtractionAtmospheric  
Oxidation Without BurningBurning and  
Subsequent Washing of Cupreous Iron  
PyritesExtraction of Copper from Burnt  
Cupreous ...

## **How to Remove Arsenic, Antimony & Bismuth from Copper**

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Alkali metal and alkaline earth arsenides. The group 1 alkali metals and the group 2, alkaline earth metals, form arsenides with isolated arsenic atoms. They form upon heating arsenic powder with excess sodium gives sodium arsenide ( $\text{Na}_3\text{As}$ ). The structure of  $\text{Na}_3\text{As}$  is complex with unusually short Na-Na distances of 328–330 pm

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which are shorter than in sodium metal.

### **Arsenide - Wikipedia**

Group 2 cations are divided into two sub-groups. Group II A cations: Hg 2+, Pb 2+, Bi 3+, Cu 2+, Cd 2+ (Copper sub-group) Group II B cations: Sn 2+, Sn 3+, As 3+ (Arsenic sub-group) Reagent for precipitating group II cations: Hydrogen

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sulfide ( $H_2S$ ) in acidic medium.

### **Answered: Explain the theoretical aspect of the... | bartleby**

Make the solution acidic by adding one or more drops of 6 M HCl. Add 1 mL of thioacetamide and stir well. Heat the test tube in the boiling water bath for 5 minutes. If antimony is present, a red

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orange precipitate of antimony sulfide should form. This same test will also work with arsenic(III), tin(II), and tin(IV).

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